A global (volume averaged) model of a chlorine discharge: Dilution with argon and oxygen

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Outline

- Chlorine is an electronegative diatomic gas that is widely used in plasma etching of both semiconductors and metals, in particular poly-silicon gate and aluminum interconnects
- Chlorine atoms are believed to be the primary reactant in plasma etching
- The chlorine molecule has
  - a low dissociation energy (2.5 eV)
  - a near-zero threshold energy for dissociative attachment
- All electronic excitations of the molecule appear to be dissociative, and no metastable molecular states are of importance

Outline

- The global (volume averaged) model
  - Model parameters
- Comparison with measurements
- Particle densities
- Sensitivity analysis
- Argon dilution
- Oxygen dilution
- Summary

The global (volume averaged) model

The global (volume averaged) model

- A steady state global (volume averaged) model was developed for the chlorine discharge using a revised reaction set
- The following species are included
  - electrons
  - the ground state chlorine molecule $Cl_2(X\,^1\Sigma_g^+, v = 0)$,
  - the vibrationally excited ground state chlorine molecules $Cl_2(X\,^1\Sigma_g^+, v = 1 - 3)$
  - the ground state chlorine atom $Cl(3p^5\,^2P)$
  - the negative chlorine ion Cl^-
  - the positive chlorine ions Cl^+ and Cl_2^+
- The content of the chamber is assumed to be nearly spatially uniform and the power is deposited uniformly into the plasma bulk

The global (volume averaged) model

- The particle balance equation for a species X is given

$$\frac{\mathrm{d}n^{(X)}}{\mathrm{d}t} = 0 = \sum_i R^{(X)}_{\mathrm{Generation},i} - \sum_i R^{(X)}_{\mathrm{Loss},i}$$

  where $R^{(X)}_{\mathrm{Generation},i}$ and $R^{(X)}_{\mathrm{Loss},i}$, respectively, are the reaction rates of the various generation and loss processes of the species X

- The power balance equation, which equates the absorbed power P_{abs} to power losses due to elastic and inelastic collisions and losses due to charged particle flow to the walls is given as

$$\frac{1}{V}\left[P_{\mathrm{abs}} - eVn_{\mathrm{e}} \sum_{\alpha} n^{(\alpha)} \mathcal{E}_{\mathrm{c}}^{(\alpha)} k_{\mathrm{iz}}^{(\alpha)} - eu_{\mathrm{B0}} n_{\mathrm{i}} A_{\mathrm{eff}} (\mathcal{E}_{\mathrm{i}} + \mathcal{E}_{\mathrm{e}})\right] = 0$$

The global (volume averaged) model

- For the edge-to-center positive ion density ratio we use

$$h_L \simeq \left[\left(\frac{0.86}{(3+\eta L/2\lambda_{\rm i})^{1/2}}\frac{1}{1+\alpha_0}\right)^2 + h_{\rm c}^2\right]^{1/2}$$

$$h_R \simeq \left[\left(\frac{0.8}{(4+\eta R/\lambda_{\rm i})^{1/2}}\frac{1}{1+\alpha_0}\right)^2 + h_{\rm c}^2\right]^{1/2}$$

where $\alpha_0 \approx (3/2)\alpha$ is the central electronegativity, $\eta = 2T_+/(T_+ + T_-)$ and

$$h_{\rm c} \simeq \left[\gamma_-^{1/2} + \gamma_+^{1/2}[n_*^{1/2} n_+/n_-^{3/2}]\right]^{-1} \quad \text{and} \quad n_* = \frac{15}{56}\frac{\eta^2}{k_{\rm rec}\lambda_{\rm i}} v_{\rm i}$$

is based on a one-region flat topped electronegative profile

$$\gamma_- = \mathrm{T_e}/\mathrm{T_-} \quad \text{and} \quad \gamma_+ = \mathrm{T_e}/\mathrm{T_+}$$

Kim et al., J. Vac. Sci. Technol. A, **24** 2025 (2006)

The global (volume averaged) model

- The diffusional losses of the neutral chlorine atoms to the reactor walls are given by

$$k_{\mathrm{Cl,wall}} = \left[\frac{\Lambda_{\mathrm{Cl}}^2}{D_{\mathrm{Cl}}} + \frac{2V(2-\gamma_{\mathrm{rec}})}{Av_{\mathrm{Cl}}\gamma_{\mathrm{rec}}}\right]^{-1} \quad \mathrm{s}^{-1}$$

  - D_{Cl} is the diffusion coefficient for neutral chlorine atoms
  - $v_{\mathrm{Cl}} = (8eT_{\mathrm{g}}/\pi m_{\mathrm{Cl}})^{1/2}$ is the mean Cl velocity
  - γ_{rec} is the wall recombination coefficient for neutral chlorine atoms on the wall surface
  - Λ_{Cl} is the effective diffusion length of neutral chlorine atoms

$$\Lambda_{\mathrm{Cl}} = \left[\left(\frac{\pi}{L}\right)^2 + \left(\frac{2.405}{R}\right)^2\right]^{-1/2}$$

- The wall recombination coefficient γ_{rec} is one of the most important parameters in chlorine discharge modelling

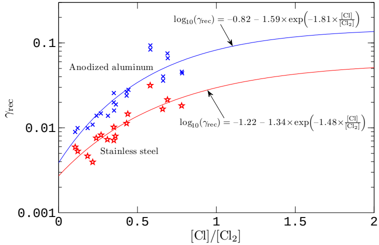
Model parameters

Surface recombination

- The wall recombination probability, $\gamma_{\rm rec}$, is a very important quantity in all low pressure molecular discharges
- We use the wall recombination coefficient measured by Stafford et al. (2009) for stainless steel

Guha et al. J. Appl. Phys., **103** 013306 (2008)

Stafford et al. J. Phys. D: Appl. Phys. **42** 055206 (2009)



A fit to the measured data is for anodized aluminum

$$\log_{10}(\gamma_{\rm rec}) = -0.82 - 1.59\ \exp\left(-1.81\times\frac{[\rm Cl]}{[\rm Cl_2]}\right)$$

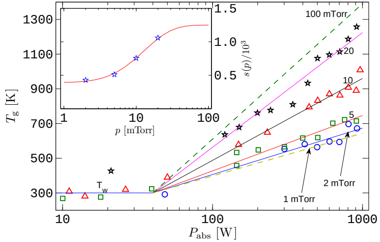
and for stainless steel

$$\log_{10}(\gamma_{\rm rec}) = -1.22 - 1.34\ \exp\left(-1.48\times\frac{[\rm Cl]}{[\rm Cl_2]}\right)$$

Gas temperature

- Donnelly and Malyshev (2000) found that the neutral chlorine gas temperature was between 300 and 1250 K, increasing with power and pressure up to 1000 W and 20 mTorr

Donnelly and Malyshev, Appl. Phys. Lett. **77** 2467 (2000)



A fit through the measured data gives

$$T_g(P_{abs}, p) = 300 + s(p) \frac{\log_{10}(P_{abs}/40)}{\log_{10}(40)}$$

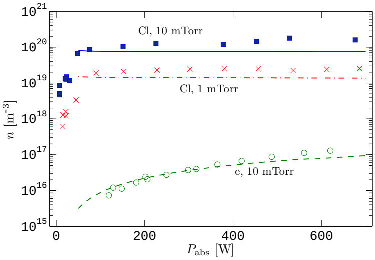
where

$$s(p) = 1250\,(1 - e^{-0.091 \times p}) + 400\, e^{-0.337 \times p}$$

Comparison with experiments

Comparison with experiments

- Densities of neutral Cl atoms and electrons versus power
- The agreement with the measured electron density is excellent
- The calculated density of atomic chlorine is in a very good agreement with the measured data at both 1 and 10 mTorr



- inductively coupled cylindrical stainless steel chamber
- $L = 20$ cm and $R = 18.5$ cm

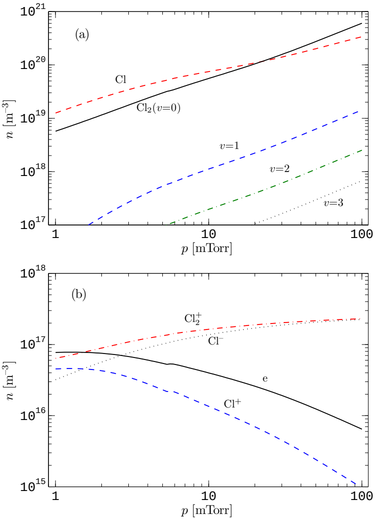
Malyshev and Donnelly, J. Appl. Phys. **88** 6207 (2000)

Malyshev and Donnelly, J. Appl. Phys. **90** 1130 (2001)

Particle densities

Particle densities

- Atomic chlorine Cl is the dominant particle at low pressure, but the chlorine molecule Cl_2 has a larger density above 20 mTorr
- The density of the atomic ion Cl^+ is always much smaller than the Cl_2^+ density, decreasing with pressure



- a cylindrical stainless steel chamber
  radius $R = 18.5$ cm
  length $L = 20$ cm

  $P_{\text{abs}} = 323$ W

Sensitivity analysis

Sensitivity analysis – EEDF

- The discharge pressure was 10 mTorr and the absorbed power 323 W
- We allow the electron energy distribution function to vary according to the general distribution function

$$f(\mathcal{E}) = c_1 \mathcal{E}^{1/2} \exp(-c_2 \mathcal{E}^x)$$

where the coefficients c_1 and c_2 depend on the energy $\mathcal{E}$ and the distribution parameter x

	$[\mathrm{Cl}]/n_\mathrm{g}$	$[\mathrm{Cl}^+]/n_+$	α	$\mathrm{T_e}$	n_e
x: 1 – 2	$\downarrow$ 1.01	$\downarrow$ 1.40	$\uparrow$ 1.34	$\uparrow$ 1.43	$\downarrow$ 1.65

Gudmundsson, Plasma Sources Sci. Technol., **10** 76 (2001)

Sensitivity analysis – γ_{rec}

	[Cl]/n_{g}	[Cl^{+}]/n_{+}	α	$\mathrm{T_e}$	n_{e}
γ_{rec}: 10^{-4} – 1	$\downarrow$ 5.75	$\downarrow$ 34.6	$\uparrow$ 4.25	$\uparrow$ 1.13	$\downarrow$ 1.59

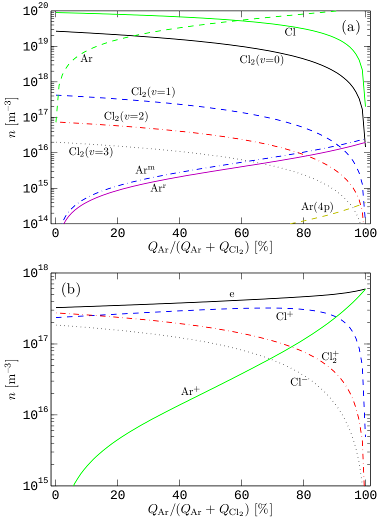
- The wall recombination coefficient γ_{rec} determines the rate coefficient for recombination of neutrals on the wall
- However, varying γ_{rec} has a much larger effect on the atomic ion fraction than on the dissociation fraction

Argon dilution

Argon dilution – particle densities

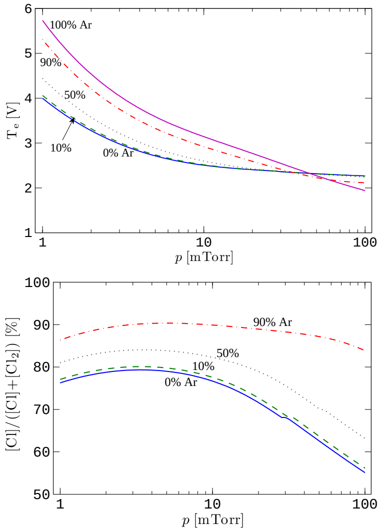
- The discharge is highly dissociated with Cl atoms being the dominant neutral until the argon content is 60%
- The Cl^+ density increases until the argon dilution is 68%
- This is likely a result of the increased electron temperature
- A cylindrical stainless steel chamber $L = 10$ cm and $R = 10$ cm
  Pressure is $p = 10$ mTorr
  Power $P_{abs} = 500$ W

  $Q_{Cl_2} + Q_{Ar} = 100$ sccm



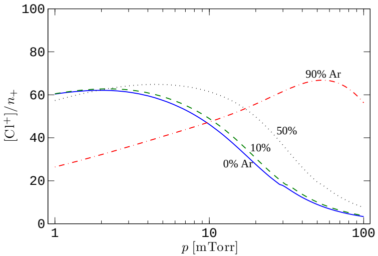
Particle densities

- The electron temperature increases with argon content at low and intermediate pressures
- The chlorine dissociation increases with argon content
- The chlorine dissociation fraction decreases with increased pressure above 10 mTorr at low and moderate argon contents



Particle densities

- The pressure dependence of the fraction of Cl^+ positive ions can be modified by argon dilution
- It peaks at low pressure when the argon content is low or moderate, but at high pressure in an argon dominated discharge
- The peak value increases slightly with increased argon content, even when the argon content has reached 90%



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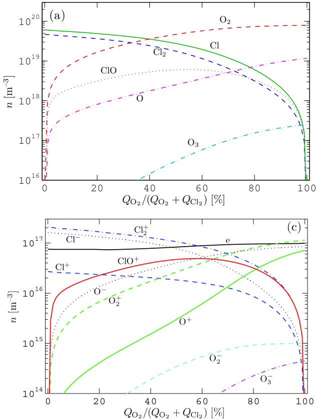
J. Phys. D., **43** 115201 (2010)

Oxygen dilution

Particle densities

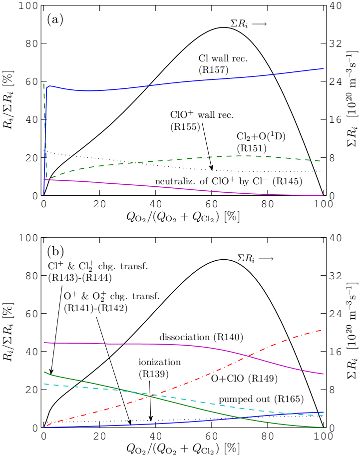
- The Cl^+ density decreases with increased oxygen dilution
- The chlorine-oxide molecule ClO and its ion ClO^+ peak when Cl_2 and O_2 flowrates are roughly equal
- The $O_2(a^1\Delta_g)$ density is about 9 – 10 % of the total O_2 density
- The electron density increases about 30 % between pure chlorine and pure oxygen discharge



A cylindrical stainless steel chamber
$L = 10$ cm and $R = 10$ cm
$p = 10$ mTorr and $P_{\rm abs} = 500$ W

Particle densities

- The total rate for creation and loss of ClO molecules is at maximum when the oxygen content is 65%.
- Wall recombination of Cl molecules, is the dominating pathway for creation of ClO molecules
- The bulk processes and recombination of ClO^+ ions at the wall account for roughly 33–43% of the total rate for ClO creation, combined



Summary

Summary

- A global model of Cl_2, Cl_2/Ar and Cl_2/O_2 discharges has been developed
- The chlorine discharge remains highly dissociated in all conditions, being over 20 % at the lowest power and highest pressure explored
- Cl^- ions are essentially entirely produced in dissociative attachment of electrons to Cl_2 and lost to mutual neutralization with Cl^+ and Cl_2^+
- The effect of vibrationally excited chlorine molecules $Cl_2(v > 0)$ is not great, at most increasing the Cl^- production by about 14 %
- The Cl^+ density increases with increased argon dilution but decreases with increased oxygen dilution
- The ClO molecule is mainly created by recombination at the discharge wall

References

Download the slides at

http://www.raunvis.hi.is/∼tumi/plasma.html

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